

Inventor: PICKFORD ET AL  
Serial No.: 10/501,538  
Filing Date: 7/16/2004  
Examiner: Christopher D. Prone  
Group Art Unit: 3738

#### REMARKS

The Office action of July 28, 2008 has been carefully considered and the application has been amended accordingly.

Previously presented claims have been cancelled. Claims 41-43 have been deleted without prejudice to presentation in a divisional application.

New claims 44-52 are submitted for consideration.

The Examiner has rejected previously presented claims under 35 U.S.C. 102 and 103 in view of U.S. Patent 5,185,075 (Rosenberg et al.) Rosenberg describes a process for anodizing a titanium implant using a phosphoric acid solution (primarily non-aqueous), and in which a small amount of silver nitrate solution may be added to precipitate any halides (which have a detrimental effect on the anodizing). The Examiner has now stated that: "The applicant argues that there is a large difference in the concentrations of silver ions absorbed between the current application and the Rosenberg patent. However this is not convincing because the claimed materials are still provided in the Rosenberg reference, which would then **inherently have the same effects** as that in the current application."

The Rosenberg reference does indeed mention anodizing, phosphoric acid, and silver nitrate. But, the overall process is a completely different process to that of the present invention; thus, the end product, as claimed, is significantly and necessarily structurally different. It differs in the manner in which the anodizing is carried out, and so in the nature of the resulting anodized surface; and it differs in the way in which silver ions are absorbed, and so in the concentration of silver ions in

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the resulting product. The differences are so significant that Rosenberg's implants would have no biocidal properties whatsoever, whereas the implants of the present invention have a significant and substantial biocidal property.

Considering first the anodizing process, Rosenberg carries out anodizing in such a way as to generate a dielectric film of titanium dioxide with high dielectric strength and with a high breakdown potential (column 3, lines 16-25), this requiring use of a substantially non-aqueous organic solvent for the phosphoric acid (column 3, lines 37-39). That is to say the anodizing is carried out in such a way as to ensure a good quality oxide layer is formed. The process typically involves a comparatively high voltage (the examples are from 100 to 500 V), and there is no point in going beyond the stage at which "age down" ends, beyond which the oxide film no longer becomes more resistive; subsequently the film integrity is attacked (column 7, lines 43-49). The resulting oxide layer is a good quality dielectric film, primarily consisting of titanium dioxide, without significant porosity or ion exchange properties.

The present invention, by contrast, carries out anodizing with an aqueous electrolyte, in such a way as to form a surface coating in the form of a matrix of titanium phosphate and/or titanium oxide which has significant ion exchange properties. The surface coating must have a degree of porosity, as a smooth surface cannot absorb sufficient silver. (If silver ions were to be absorbed as a monolayer onto a smooth surface, over the entire surface, the absorption would be about 0.3  $\mu\text{g}/\text{cm}^2$ ; in contrast the silver level measured in the present invention is 73  $\mu\text{g}/\text{cm}^2$ .)

As explained in the previous submission, i.e., in the Amendment filed on December 18, 2006, in Rosenberg's process a

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small amount of silver nitrate may be added to the phosphoric acid, to ensure that any halides are precipitated; sufficient silver nitrate may be added that (yellow) silver phosphate is precipitated. But, it will be appreciated by those having ordinary skill in this art that the particles of silver halide and silver phosphate that precipitate are not part of the surface oxide layer. Furthermore, the concentration of silver ions in solution must be very low, because silver phosphate is substantially insoluble. Since the solution is acidic, any hypothetical ion exchange of silver ions with the titanium dioxide is suppressed by the much greater concentration of hydrogen ions.

In contrast, in the process of the present invention, after anodizing, the implant is rinsed; only then is it contacted with an aqueous solution of 0.1 M silver nitrate. So, the ion exchange takes place using a liquid in which the silver ion concentration is about 1,000,000 times greater than the hydrogen ion concentration (pH7 means the hydrogen ion concentration is  $10^{-7}$  M = 0.1  $\mu$ M). Hence, the ion exchange process - considering only the difference in the solutions - is about 500,000,000 times more effective in the present invention than in the Rosenberg process.

Taking both these considerations into account, the amount of silver that can be absorbed by ion exchange into an implant of the present invention, as compared to one treated by Rosenberg's process, will be considerably greater than 500,000,000 times greater!

U.S. Patent 6,113,636 (Ogle) points out that: "Healthy humans generally have for example **plasma levels of Ag of about 0.2  $\mu$ g/L to about 10  $\mu$ g/L**, where 10  $\mu$ g/L corresponds to about 10 ppb ... Serum silver ion concentrations of about 300 ppb have

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been associated with toxic symptoms ... Silver ion concentrations of about 4 mg/L [= 4000 µg/L] are known to cause rapid cell death. Therefore, it is preferable to keep silver ... ion concentrations in the bloodstream safely below toxic levels and preferably below levels where any symptoms are observable." (Column 5, lines 10-13 and 21-28).

That information would appear to be objective and uncontentious, at least as regards the broad picture. For an implant to have any biocidal properties, it clearly must contain sufficient silver ions to raise the levels of Ag significantly above those ordinarily present. But, the range between the lowest normal level in a healthy human (0.2 µg/L) and concentrations that could be fatal (say 4000 µg/L) differ by only a factor of 20,000. The implants of the present invention do have such a biocidal effect. Assuming that the implants of the present invention generate a silver ion concentration in the plasma of say 100 µg/L, then the corresponding silver ion concentration provided by Rosenberg's procedure (taking only the difference in the solutions into account) would be less than 0.2 µg/L, that is to say about a million times less than the lowest normal level in human plasma, and so clearly having no significant effect.

It is thus clear that Rosenberg's procedure produces an implant with a much less porous oxide surface, with markedly less ion exchange properties; and that the extent to which silver is absorbed by ion exchange is completely insignificant and negligible. Such an implant would produce no change in the level of biocidal ions already present in the patient! And so it "inherently" does not have the effect of the present invention, as claimed.

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By contrast, the present invention provides an implant with a significantly more porous oxide surface, with considerably greater ion exchange properties; and silver is absorbed by ion exchange by a far more effective procedure, leading to an enormously greater loading of silver. Hence, the implants of the present invention have a significant biocidal effect.

Accordingly, applicants submit that the present claims clearly distinguish from Rosenberg et al. in the quantity of biocidal metal ions being such that the biocidal material is effective in suppressing infection after the surgical procedure. Accordingly, applicant's believe that that claims 44-52 are directed to allowable subject matter and such action by the Examiner is courteously solicited.

Favorable action on the merits is courteously solicited.

Respectfully submitted,

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Date

  
Signature